## CML 100: 2017-2018 Quantum Tutorial 3

- 1. Identify  $\widehat{H}^{(0)}$ ,  $\widehat{H}^{(1)}$  and ground state  $\psi^{(0)}$  and  $E^{(0)}$  for the following systems
  - a. An oscillator with potential

$$V(x) = \frac{1}{2}kx^2 + \frac{1}{6}\gamma x^3 + \frac{b}{24}x^4$$

b. A particle constrained to move in the region 0 < x < a with the potential

$$V(x) = 0 \quad 0 \le x < \frac{a}{2}$$
$$= b \quad \frac{a}{2} < x < a$$

- c. A Helium atom
- d. A H-atom in an electric field of strength F. Hamiltonian for this system is

$$\widehat{H} = -\frac{\hbar^2}{2m_e} \nabla^2 - \frac{e^2}{4\pi\varepsilon_0 r} + eFr\cos\theta$$

e. A rigid rotor with a dipole moment  $\mu$  in an electric field of strength F. Hamiltonian is

$$\widehat{H} = -\frac{\hbar^2}{2I} \nabla^2 + \mu F \cos \theta$$

2. Using the P-I-B as the unperturbed problem, calculate the first order correction to the ground state energy for a particle constrained to move in the region 0 < x < a with the potential

$$V(x) = 0 \quad 0 \le x < \frac{a}{2}$$
$$= b \quad \frac{a}{2} < x < a$$

- 3. If you use the trial function  $\phi(x) = (1 + cax^2)e^{-ax^2/2}$  where  $a = (k\mu/\hbar^2)$  and c is the variational parameter for the ground state energy of a Harmonic Oscillator, what do you expect the value of c should be? Why?
- 4. The Hamiltonian for the 3D Harmonic oscillator can be written as

$$\widehat{H} = -\frac{\hbar^2}{2\mu r^2} \frac{d}{dr} \left( r^2 \frac{d}{dr} \right) + \frac{1}{2} k r^2$$

Use the trial wavefunctions  $e^{-\alpha r^2}$  and  $e^{-\alpha r}$  and suggest which one is better and why.

- 5. Apply the first order perturbation correction to He atom. Compare the results with the variational results from the lecture.
- 6. The Morse potential  $V(x) = D(1 e^{-\beta x})^2$  is a good approximation to the intermolecular potential energy for a diatomic molecule. The constants D and  $\beta$  for  $H_2$  molecule are  $D = 7.61 \times 10^{-19}$  J and  $\beta = 0.0193$  pm<sup>-1</sup>. Use the expansion of the exponential to see the similarity with the Anharmonic oscillator. Relate the constants D and  $\beta$  with those of the harmonic oscillator.

- 7. The P-I-B problem is fully solvable as you have seen. However, use the trial wavefunction  $\phi = \sum_{n=1}^{2} c_n x^n (a-x)^n$  to perform variational calculation for the ground state and compare the results. Is the trial function well-behaved?
- 8. The spin operators also obey the commutation relations

$$\left[\hat{S}_{x},\hat{S}_{y}\right]=i\hbar\hat{S}_{z}\;;\;\left[\hat{S}_{y},\hat{S}_{z}\right]=i\hbar\hat{S}_{x}\;,;\;\left[\hat{S}_{z},\hat{S}_{x}\right]=i\hbar\hat{S}_{y}$$

We define  $\hat{S}_+ = \hat{S}_x + i\hat{S}_y$  and  $\hat{S}_- = \hat{S}_x - \hat{S}_y$ .

- (a) Show that  $\hat{S}_{+}$  and  $\hat{S}_{-}$  are non-Hermitian.
- (b) Show that  $[\hat{S}_z, \hat{S}_+] = \hbar \hat{S}_+$  and  $[\hat{S}_z, \hat{S}_-] = -\hbar \hat{S}_-$
- (c) Show that  $\hat{S}_{+}\hat{S}_{-} = \hat{S}^{2} \hat{S}_{z}^{2} + \hbar \hat{S}_{z}$  and  $\hat{S}_{-}\hat{S}_{+} = \hat{S}^{2} \hat{S}_{z}^{2} \hbar \hat{S}_{z}$ Where,  $\hat{S}^{2} = \hat{S}_{x}^{2} + \hat{S}_{y}^{2} + \hat{S}_{z}^{2}$
- (d) What is the result of operating  $\hat{S}_{+}$  and  $\hat{S}_{-}$  on the two spins  $\alpha$  and  $\beta$ ?
- 9. The total z component of the spin angular momentum operator for an N-electron system is

$$\hat{S}_{z,total} = \sum_{j=1}^{N} \hat{S}_{z,j}$$

Show that

$$\psi = \frac{1}{\sqrt{3!}} \begin{vmatrix} 1s\alpha(1) & 1s\beta(1) & 2s\alpha(1) \\ 1s\alpha(2) & 1s\beta(2) & 2s\alpha(2) \\ 1s\alpha(3) & 1s\beta(3) & 2s\alpha(3) \end{vmatrix}$$

is an eigenfunction of  $\hat{S}_{z,total}$ .

- 10. Use the z-axis as the internuclear axis and evaluate the overlap integral for an 1s-orbital on one H-atom and a  $2p_x$  orbital on the other H-atom. (Hint: You need not solve any integrals).
- 11. Determine the M.O.'s for 1,3-butadiene using the Hückel MOT. Sketch the MOs showing the size of the p-orbitals on the four carbon atoms proportional to the coefficients.
- 12. Calculate the delocalization energy for benzene and M.O.s. Sketch the M.O.s.